

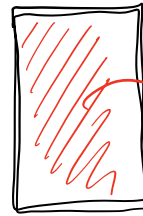
EX06 (fiche de préparation du TE du 05/11/2025)

$$10^6 \mu g = 1 g$$

micro

$$20 l \text{ à } 0,955 \text{ kg/dm}^3$$

23,2 kg masse TOTALE



20l

$$20 l = 20 \text{ dm}^3$$

$$1 l = 1 \text{ dm}^3$$

$$20 \text{ dm}^3 \cdot 0,955 \text{ kg/dm}^3$$



V



ρ

$$= 19,1 \text{ kg}$$

$$\rho = \frac{m}{V}$$

$$m = \rho \cdot V = V \cdot \rho$$

En tout

$$23,2 \text{ kg}$$

masse du m3zont

$$\Rightarrow 23,2 - 19,1 = \boxed{4,1 \text{ kg}} \quad \text{masse du jerrycan}$$

$$\boxed{10^6 \mu\text{g} = 1\text{g}}$$

$$\boxed{1\text{kg} = 1000\text{g}}$$

$$4,1 \text{ kg} = 4100 \text{ g} = 4100 \cdot 10^6 \mu\text{g}$$

$$= \boxed{4,1 \cdot 10^9 \mu\text{g}}$$

EXO 7

photo : redite'
42500 : 1

photo	redite'
9 cm	?
42500 cm	1 cm

$$? = \frac{9 \cdot 1}{42500}$$

$$\frac{9}{42500} \text{ cm} \cong 0,000211764 \text{ cm}$$
$$\cong 2,11764 \cdot 10^4 \text{ cm} \cong 211764 \cdot 10^0 \mu\text{m}$$

$$10^6 \mu\text{m} = 1\text{m} = 100 \text{ cm}$$

$$10^4 \mu\text{m} = 1\text{cm}$$

05/11/2025

65'

Col. littéral

Not. scientifique :

2.14

⚠ Problèmes selon liste

EX08

$$18 \text{ g/mol} = 18 \text{ g/mole}$$

$$6,022 \cdot 10^{23} \text{ molecules/mole}$$

$$1 \text{ g H}_2\text{O} \leftrightarrow 1 \text{ cm}^3 \text{ H}_2\text{O}$$

$$1 \text{ dm}^3 = 1000 \text{ cm}^3 = 1000 \text{ g}$$

$$\frac{1000 \text{ g}}{18 \text{ g/mole}} = \frac{500}{9} \text{ moles}$$

$$= 55,55 \text{ moles}$$

$$\# \text{ molecules: } 6,022 \cdot 10^{23} \cdot 55,55$$

$$334,55 \cdot 10^{23} = 3,346 \cdot 10^{25}$$